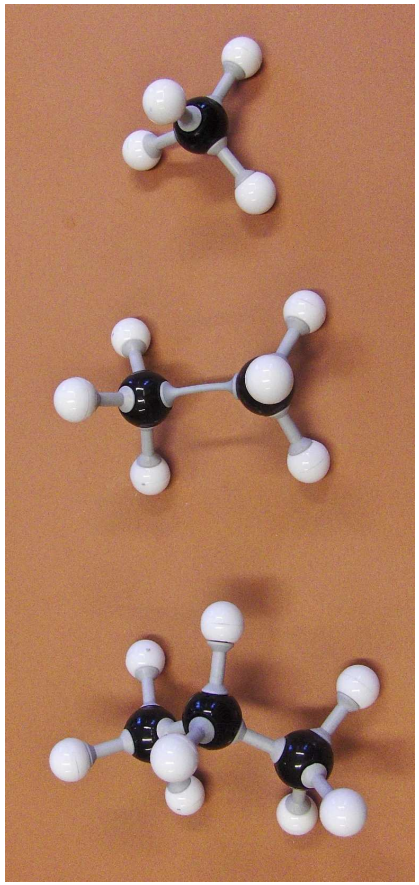
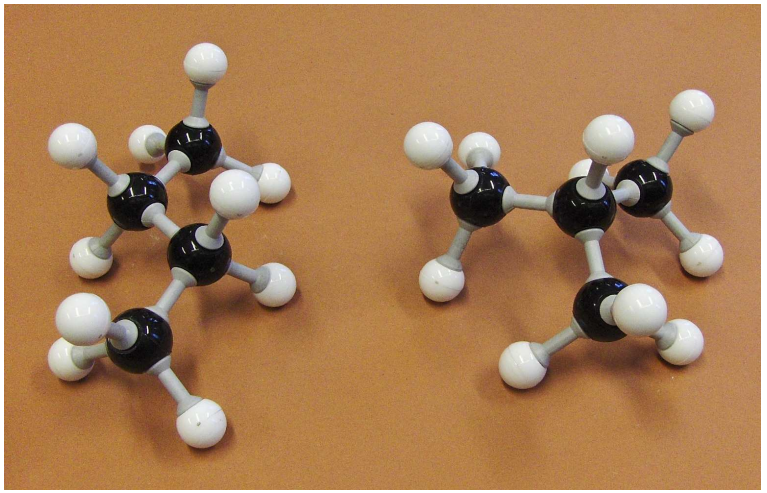
Gas

Basics

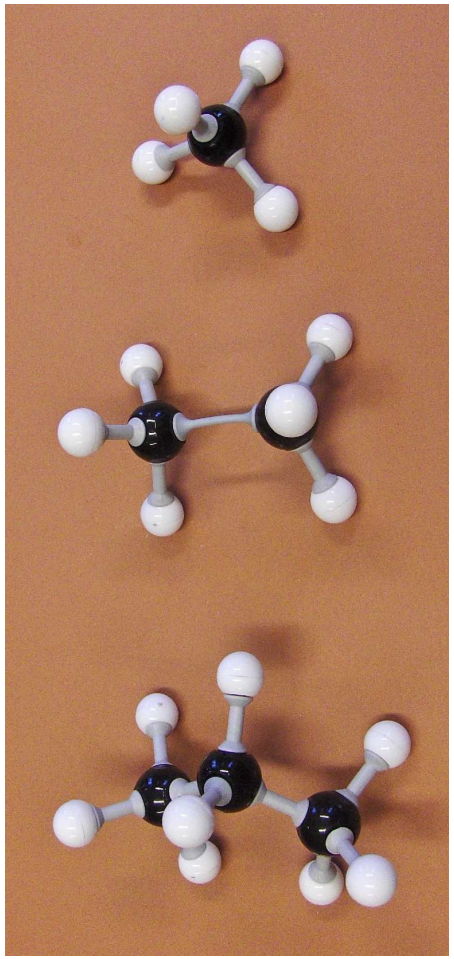
Rich gas from fields



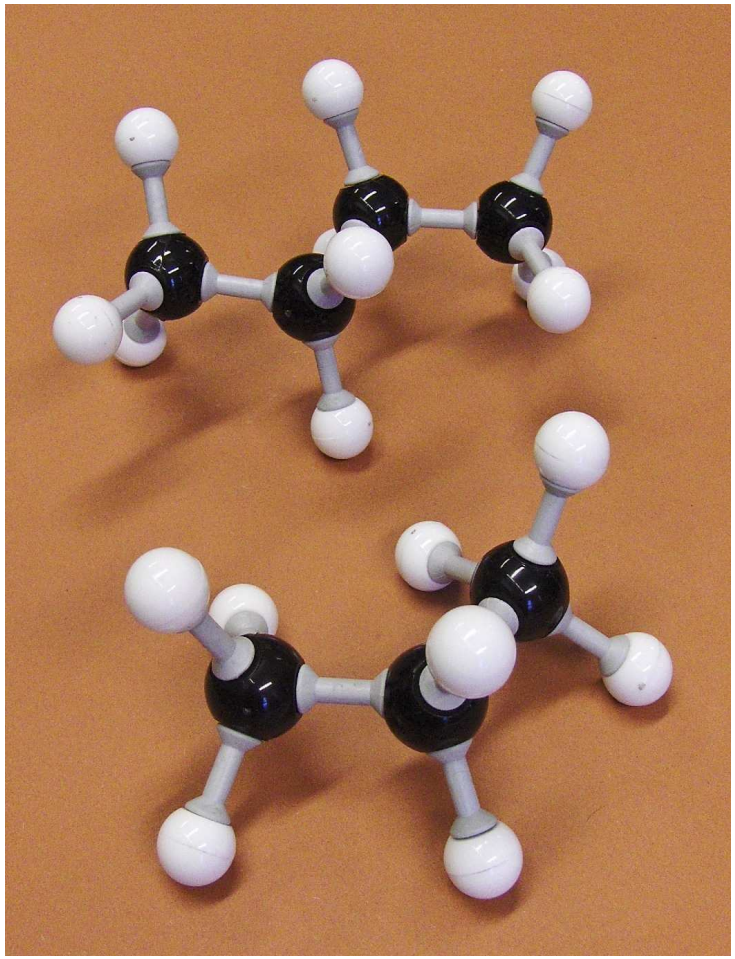
Methane, ethane, propane, n-butane, iso-butane and naphtha.



Methane – Ethane - Propane



Propane C_3 Butane C_4



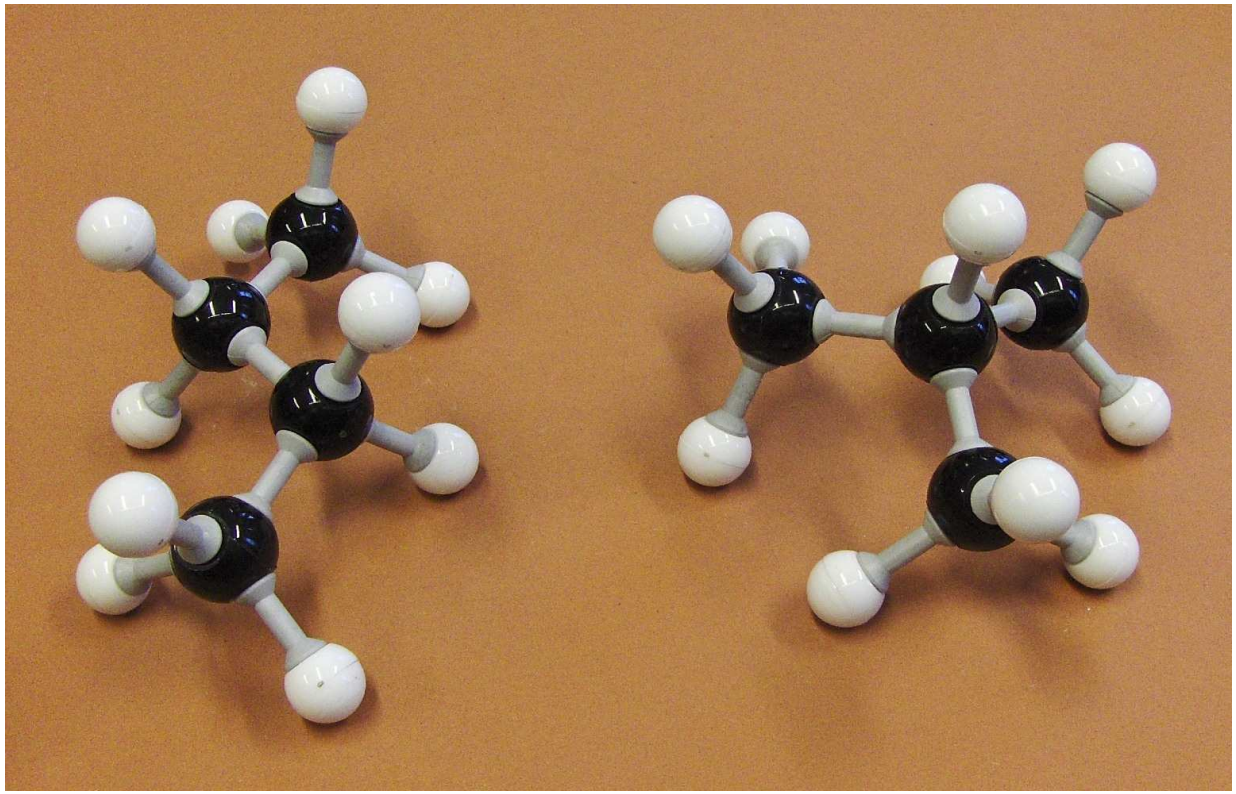
Propane: C3H8

Camping gas

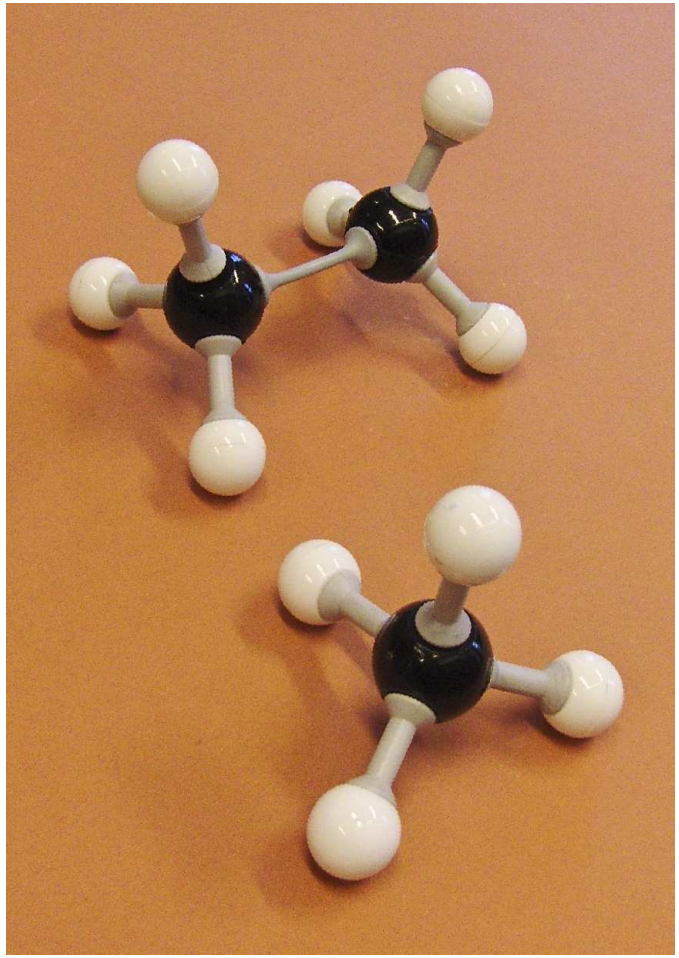
Butane: C4H10

Lighter gas and small camping burners.

n-Butan and iso-Butane

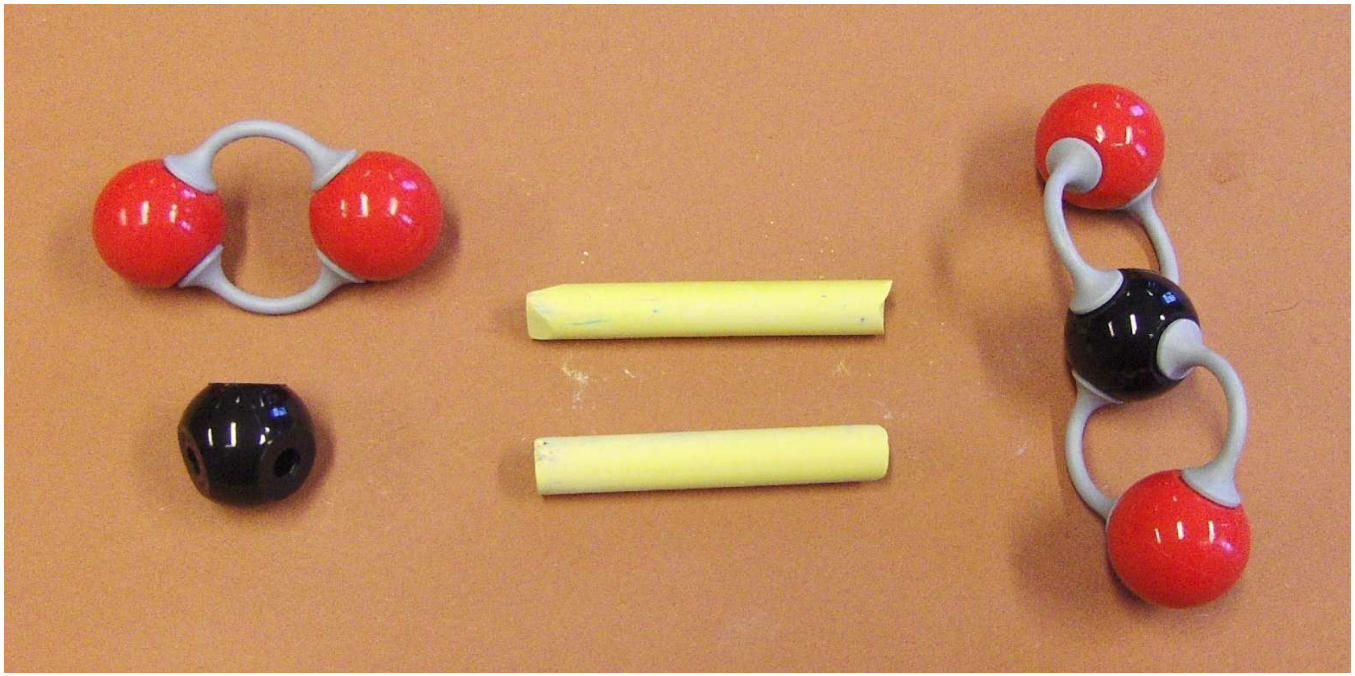


Methane - Ethane



	Parts of natural gas
Etan C_2H_4	5 – 15 %
Metan CH_4	85 – 95 %

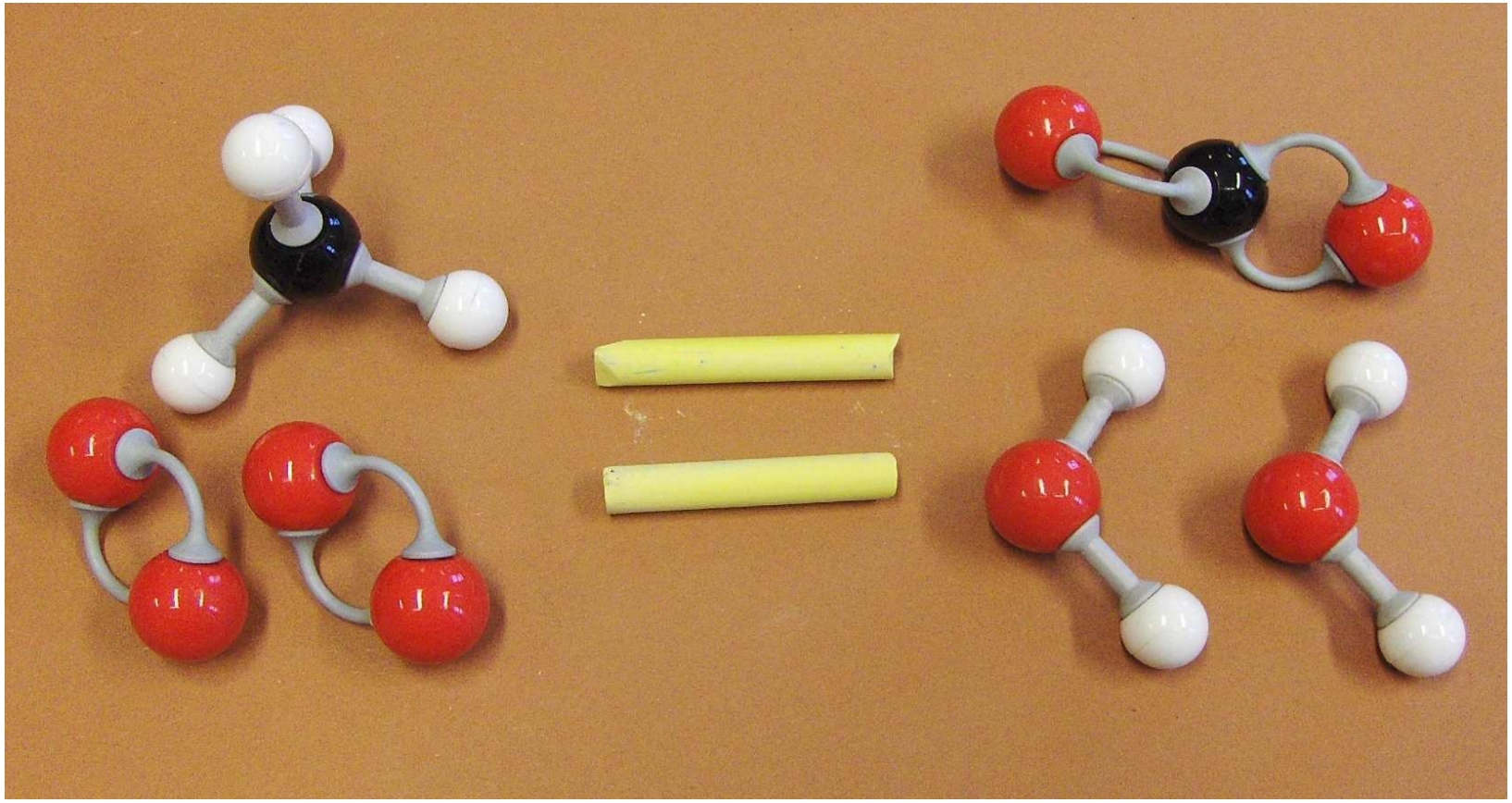
Burning of carbon and hydrocarbons
Carbon



Carbon + oxygen = carbondioxide

One atom carbon needs one molecule of oxygen and gives one molecule og carbondioxide as a result.

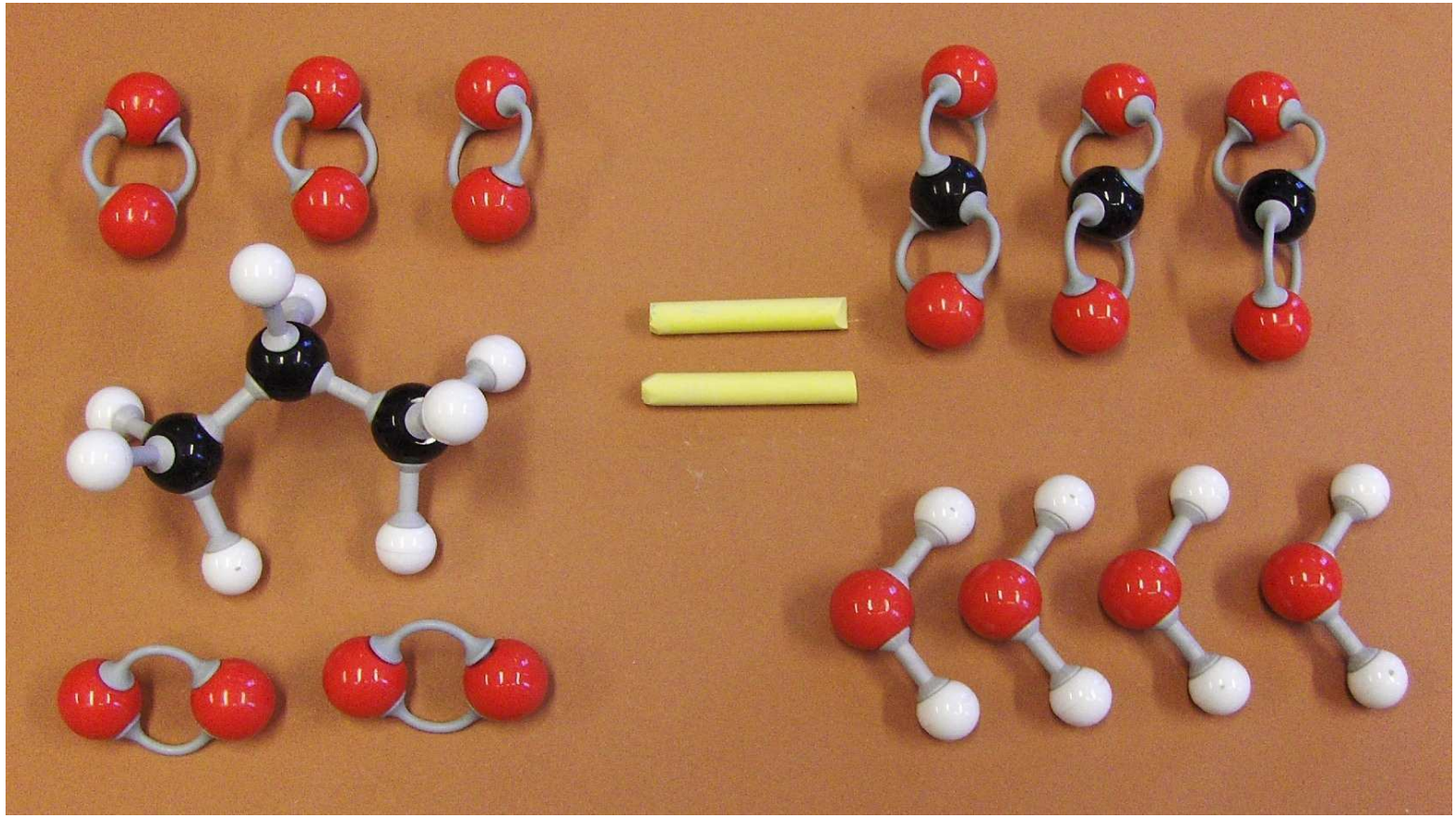
Methane burns



Methane + oxygen = water + carbondioxide

One molecule of methane needs two molecules of oxygen and gives one molecule of carbondioxide and to molecules of water as a result.

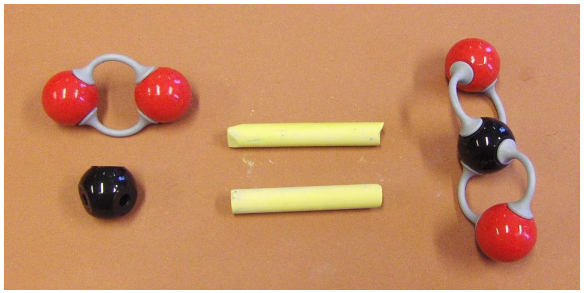
Propane burns

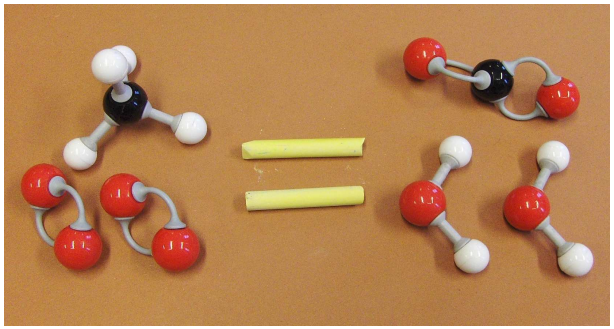


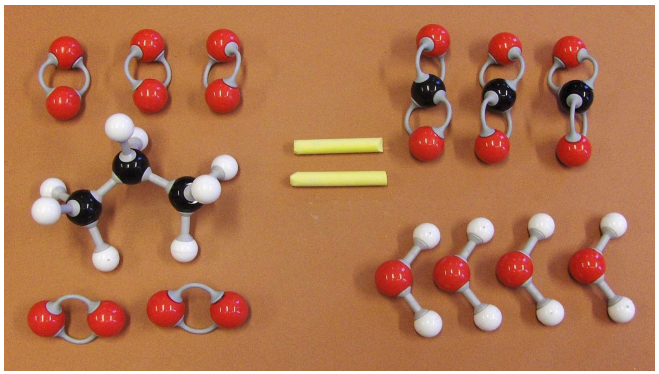
One molecule of propane needs 5 molecules of oxygen and gives 3 molecoles of carbondioxide and 4 molecules of water.

Burning of carbon and hydrocarbons

Summary







CO_2 generated	
Carbon	100 %
Methane	33 %
Propane	43 %

Lighter hydrocarbons gives more water and less CO_2 than pure carbon and heavier hydrocarbons.

The gases don't contain sulphur.